
Safety Assessment of Barium Sulfate as Used in Cosmetics

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All interested persons are provided 60 days from the above date to comment on this Tentative Report and to identify additional published data that should be included or provide unpublished data which can be made public and included. Information may be submitted without identifying the source or the trade name of the cosmetic product containing the ingredient. All unpublished data submitted to CIR will be discussed in open meetings, will be available at the CIR office for review by any interested party and may be cited in a peer-reviewed scientific journal. Please submit data, comments, or requests to the CIR Director, Dr. Lillian J. Gill.

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ABSTRACT: Barium sulfate functions as an opacifying agent in cosmetic products, and is being used at concentrations up to 0.99% and 37% in rinse-off and leave-on products, respectively. The Panel noted that the history of safe medical use of barium sulfate indicates no significant toxicity concerns relating to systemic exposure to these ingredients. Furthermore, the extensive clinical experience of the Panel, including the results of numerous patch tests, indicates that barium salts do not have the potential to induce sensitization. The Panel noted that salts of sulfuric acid, such as sodium sulfate, can be irritating to the skin, so cosmetic products containing barium sulfate should be formulated to be non-irritating. The Panel concluded that barium sulfate is safe in the present practices of use and concentration in cosmetics, when formulated to be non-irritating.

INTRODUCTION

The safety of barium sulfate, an inorganic sulfate, as used in cosmetics is reviewed in this safety assessment. Barium sulfate functions as an opacifying agent in cosmetic products.

CHEMISTRY

Definition and Structure

The inorganic sulfates (See Figure 1) are salts of sulfuric acid. Except for the ammonium salts, these ingredients are mineral salts readily found in nature (but may also be easily synthesized).¹

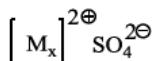


Figure 1. Inorganic Sulfates – wherein “M” is an ammonium or metal cation

Barium sulfate is the inorganic salt that conforms to the formula that is included in Figure 2 below.² It is the barium salt of sulfuric acid.



Figure 2. Formula for Barium Sulfate

Physical and Chemical Properties

Barium sulfate, available in the form of polymorphous crystals, has a molecular weight of 233.39, and is soluble in sulfuric acid, but insoluble in water.³

Method of Manufacture

The inorganic sulfates are typically manufactured by mining of natural minerals (as many inorganic sulfates occur naturally in hydrated form) or by reaction of available ore or inorganic oxides, hydroxides, or carbonates, with sulfuric acid.¹ These methods produce hydrated inorganic sulfates. To produce the anhydrous salts, an additional step of dehydration (e.g., by heating and reduced pressure) must occur. For example, hydrated magnesium sulfate can be mined as kieserite or epsomite (Epsom salts), or it can be prepared by dissolving magnesium oxide, magnesium hydroxide, or magnesium carbonate in sulfuric acid. Heating of this hydrate reversibly drives off water and produces anhydrous magnesium sulfate (a potent desiccant).

Composition/Impurities

Barytes is the naturally occurring rock form of BaSO₄.⁴ A study was performed to characterize the mineralogical forms of barium and the trace heavy metal impurities in commercial barytes of different origins using electron probe microanalysis (EPMA), X-ray diffraction (XRD), and inductively coupled plasma mass spectrometry (ICP-MS). Qualitative EPMA results indicated the presence of different minerals in commercial barytes, including barite (BaSO₄), barium feldspar, galena (PbS), pyrite (FeS₂), sphalerite (ZnS), quartz (SiO₂), and silicates. Quantitative EPMA confirmed that the barite

crystals in the barytes contain some strontium and a little calcium, whereas, trace heavy metals occur in the associated minerals. Analysis of *aqua regia* extracts of barytes samples by ICP-MS has indicated the presence of a large number of elements in the associated minerals. Arsenic, copper, and zinc concentrations correlate closely in all 10 samples.

Chromium has been detected in commercial samples of pharmaceutical grade barium sulfate at concentrations ranging from 0.45 to 1.06 µg/g.⁵

USE

Cosmetic

Barium Sulfate functions as an opacifying agent in cosmetic products.²

According to information supplied to the Food and Drug Administration (FDA) by industry as part of the Voluntary Cosmetic Registration Program (VCRP) in 2014, barium sulfate is being used mostly in leave-on products.⁶ Results from a survey of ingredient use concentrations provided by the Personal Care Products Council in 2013 indicate that this ingredient is being used at concentrations up to 0.99% and 37% in rinse-off and leave-on products, respectively.⁷ Summarized data on frequency and concentration of use in cosmetics are presented in Table 1.

Cosmetic products containing barium sulfate may be applied to the skin and nails, or, incidentally, may come in contact with the eyes and mucous membranes. Products containing these ingredients may be applied as frequently as several times per day and may come in contact with the skin or hair for variable periods following application. Daily or occasional use may extend over many years.

Barium sulfate is being used at concentrations up to 15% in cosmetic products that are sprayed (perfumes) and at concentrations up to 15.8 % in powders (face and dusting powders). Because this ingredient is used in aerosol/pump hair sprays and in powders, it could possibly be inhaled. In practice, 95% to 99% of the droplets/particles released from cosmetic sprays have aerodynamic equivalent diameters >10 µm, with propellant sprays yielding a greater fraction of droplets/particles below 10 µm, compared with pump spray.^{8,9,10,11} Therefore, most droplets/particles incidentally inhaled from cosmetic sprays would be deposited in the nasopharyngeal and bronchial regions and would not be respirable (i.e., they would not enter the lungs) to any appreciable amount.^{8,9}

Non-Cosmetic

Barium sulfate has been used as a gastrointestinal contrast agent in roentgenographic procedures.¹²

Barytes (the naturally occurring rock form of BaSO₄) has been referred to as the standard densification agent used in drilling fluids worldwide.⁴

TOXICOKINETICS

Rats inhaled barium sulfate (40 mg/m³) for up to 2 months.¹³ The rats were killed at 2-week intervals after the beginning of the exposure period, and the barium content of the lungs, lymph nodes, jaw, and femur was determined. Lymph transport was negligible. The barium content of bone increased initially, and then gradually decreased during treatment. After 2 weeks of exposure, the barium content of the lungs was high, but decreased rapidly and then increased considerably over the next 4 weeks of treatment.

Following endotracheal administration of barium sulfate (BarosperserTM) into 220 Sprague-Dawley rats and 3 dogs, radiographic and histologic studies were performed.¹⁶ BarosperserTM (0.25 ml), under fluoroscopic control, was injected endotracheally into rats, and a 1.75 ml/kg dose of BarosperserTM was administered endotracheally to dogs. The 0.25 ml dose was selected based on the results of a pilot study in which all 30 rats died after intratracheal administration of 0.5 ml BarosperserTM. After dosing with BarosperserTM, all of the rats and 2 dogs were radiographed for a total of 9 months. BarosperserTM was virtually cleared from the trachea and stem bronchi in 30 minutes. BarosperserTM cleared more slowly from the lungs of 3 dogs, when compared to these data on rats. Additonal study results are included in sections on Acute Intratracheal Toxicity and Inflammatory Response.

TOXICOLOGY

Acute Toxicity

Oral

Six groups of 16 to 26 young male CBL-Wistar albino rats (ages not stated) received the following 6 total doses of barium sulfate (150% w/v suspension), respectively, by intragastric cannula: 188, 225, 263, 300, 338, and 375 g/kg.¹⁴ For each total dose, 40% was given initially, followed by 35% of the dose 3 h later, and 25% 4 h later. Fifty control rats were dosed with distilled water. Fifty experimental animals died from stomach rupture, and the mean LD₅₀ was 307 ± 29 g/kg. Stomach rupture appeared to have been due, in part, to failure of the animal to pass barium sulfate along the gastrointestinal tract. In 90% of the animals, hemorrhagic areas were found in the gastric mucosa, mainly on the anterior and posterior surfaces. The adrenal glands were enlarged, the liver was small, and the stretched abdominal muscle had a watery consistency.

Intratracheal

Following endotracheal administration of barium sulfate (BarosperseTM) into 220 Sprague-Dawley rats and 3 dogs, radiographic and histologic studies were performed.¹⁶ BarosperseTM (0.25 ml), under fluoroscopic control, was injected endotracheally into rats, and a 1.75 ml/kg dose of BarosperseTM was administered endotracheally to dogs. The 0.25 ml dose was selected based on the results of a pilot study in which all 30 rats died after intratracheal administration of 0.5 ml BarosperseTM. After dosing with BarosperseTM, all of the rats and 2 dogs were radiographed for a total of 9 months. Diffuse, but patchy, acinar filling resulted in a slow decrease in barium aggregates from the lungs of rats and dogs over a 9-month period. New infiltrates were found in 15% of the rats on serial follow-up. Two of the 3 dogs dosed with BarosperseTM died during the first 24 h; both dogs had diffuse alveolarization of the contrast agent.

Inflammatory Response

Inhalation

Groups of male Wistar rats were exposed (whole-body) 5 days per week (7h/day) to barium sulfate dust.¹⁷ At each of 6 time points, 12 rats were drawn, 6 for bronchoalveolar lavage (BAL) and 6 for dust burden measurements. Of the 6 time points (all not specified) included in the protocol, only the following 4 were clearly defined: 42 days, 50 days, 70 days, and 90 days. Animals destined for bronchoalveolar lavage studies were killed 18 h after completion of the final day of exposure for that time point. Considering that dust deposited higher in the respiratory tract would have time to clear, animals used for lung and dust burden analyses were killed 66 h after the end of exposure. Three age-matched sham-exposed animals were used as controls at each time point for each test condition in the lavage studies.

The results of BAL fluid analyses indicated that: the time course of neutrophil recruitment during exposure to barium sulfate resembled that of lymph node burden; barium sulfate dust produced a low degree of inflammation at the last 3 time points of the higher (75 mg/m³) exposure concentration; the mean numbers of alveolar macrophages did not change significantly when compared to the background level in control animals; and animals exposed to barium sulfate dust had significantly higher numbers of lymphocytes in BAL fluid when compared to controls.¹⁷

Histological sections from animals killed at timepoint 6 indicated that inhalation of barium sulfate elicited accumulation of pulmonary macrophages around the dust deposition sites. The highest concentrations of macrophages with phagocytosed dust were at the bifurcations of the terminal airways and bronchioles. In some cases, there was an accumulation of inflammatory cells, including fibroblasts in the interstitium. Some macrophages with their dust burdens had become interstitialized as well, with the lesions becoming microgranulomas. In most cases where centriacinar macrophage aggregations were found, the walls of surrounding alveoli appeared thickened, mainly due to the rounding of epithelial cells, indicative of Type II cell hyperplasia. Barium sulfate did not show significant fibrogenic activity in this study.¹⁷

Subcutaneous

The effect of intrauterine, s.c. injection of sterile barium sulfate into rabbit fetuses was evaluated. Healthy pregnant rats (number not stated) of gestational periods ranging from 21 to 26 days were used.¹⁵ Two fetuses were selected at random. The dorsum of the fetus was delivered and a sterile aqueous suspension of micro-opaque barium sulfate was injected into the subcutaneous tissue of each dorsolateral surface. The hysterotomy wounds were then closed. The pregnancy was allowed to progress and fetuses were removed at varying postoperative intervals for morphological studies. Similar s.c. injections were

performed in newborn rabbits. The rabbits were killed at regular intervals and the morphology of the wounds was studied. Subpannicular injection of sterile barium sulfate in newborn rabbits produced an acute inflammatory response that was observed clearly at 24 h and well-established by 48 h. The process of repair had begun by day 4, whereby the appearance of proliferating capillaries and fibroblasts was observed. Both vascular and cellular components of the acute inflammatory response were more prominent in rabbit fetuses, and appeared earlier (well-developed within 24 h) when compared to newborn rabbits. The process of repair also began earlier; the proliferation of capillaries and fibroblasts was prominent by 48 h. By day 4, the lesion was compact, less cellular, and relatively avascular.

Endotracheal

Barium sulfate (BarosperseTM) was administered endotracheally into 220 Sprague-Dawley rats and 3 dogs.¹⁶ Barium in the alveoli and a mild acute inflammatory response were observed in rats at microscopic examination. A few hours after instillation, macrophages were observed in the alveoli and subsequently became evident in thickened septa. Focal alveolar wall granulomata were also observed. After 3 months, focal areas of acute and chronic inflammatory cells with focal fibrosis persisted, and areas of atelectasis and emphysema were also observed. At 6 months, aggregates of macrophages containing barium were the main finding. At 9 months, nodules of phagocytic cells in bronchioles and perivascular structures persisted. At 9 months after BarosperseTM instillation into the lungs, the same histological findings were observed in 3 dogs.

Case Reports

Exposure to barium sulfate occurs in miners of barium and its salts, workers in the lithopone industry, and in patients undergoing diagnostic roentgenography of the gastrointestinal tract.¹⁸ Barium sulfate dust, when inhaled, leads to a benign form of pneumoconiosis (baritosis), which occurs primarily in miners and workers in the lithopone industry. Escape of barium sulfate from the digestive tract into the peritoneal cavity has been reported in patients with peptic ulcers undergoing x-ray studies. Barium granulomas have been reported in the appendix, sigmoid and peritoneum, and rectum in patients receiving barium enemas.

A 43-year-old patient was diagnosed with acute appendicitis, and barium sulfate was used in diagnostic studies, i.e., upper gastrointestinal series and barium enema.¹⁸ During these procedures, barium sulfate entered the appendix and escaped into the mesoappendix and adjacent periappendical fat. The resulting foreign-body granuloma was said to have been due to the escape of barium sulfate.

A case of barium sulfate-granulomatosis of the lung was reported for a 67-year-old man, due to barium sulfate aspiration during an x-ray investigation of the stomach.¹⁹ In the lung parenchyma, multiple granulomas were observed in groups of alveoli where barium sulfate had been deposited.

Barium enema examination is a frequently performed radiographic procedure, and this procedure was reported to caused barium granuloma of the rectum in 2 patients (males 75 and 78 years old).²⁰ Rectal intramural extravasation of barium occurs as a result of asymmetric enema balloon inflation and impaction of the enema tip against the rectal mucosa. The lesions appeared as indurated, ulcerated rectal masses that resembled carcinoma on endoscopic examination. Deep mucosal biopsy results demonstrated no malignancy or barium sulfate crystals in tissue macrophages. Radiographs showed persistent soft-tissue barium in the rectum.

A severe anaphylactic reaction was observed in a 51-year-old female cancer patient at approximately 5 to 10 minutes after starting a barium enema.²¹ The barium enema mixture contained barium sulfate, sodium benzoate, potassium sorbate, citric acid, sodium saccharin, ethyl maltol, vegetable gum, sorbitol, simethicone, and natural and artificial flavors. It was stated that the anaphylactic reaction could have been an IgE-mediated hypersensitivity reaction to one of the barium sulfate suspension constituents. The patient had a history of prior sensitizing exposure to barium radiographic contrast material. No skin prick test reaction, i.e., no cutaneous hypersensitivity, to diluted sodium benzoate, potassium sorbate, or whole liquid barium sulfate suspension was detected. The patient declined further provocation testing. The authors noted that the patient's severe reaction to barium may have been partly attributable to the following 3 factors: (1) her history of atopy and prior medication allergy, (2) a prior sensitizing exposure to barium sulfate, and (3) possible increased absorption of allergens into the bloodstream through the recurrently bleeding ulcerated carcinoma of the sigmoid colon.

Two children developed hypersensitivity reactions of varying severity following upper GI series.²² The first case involved an 11-year-old boy with documented anaphylaxis, following exposure to fish and peanuts, and multiple food intolerances. The patient experienced oral swelling and a red swollen tongue after drinking 150 ml of 45% weight/weight barium sulfate. The absence of sequelae after prior upper GI series that involved drinking barium sulfate was noted. Endoscopic biopsies from the upper and lower GI tracts established the diagnosis of eosinophilic gastroenteropathy. The

second case involved a 7-year-old girl with a history of mild allergy to penicillin (hives), but no other known allergies. After drinking 150 ml of 45% barium sulfate, she developed urticaria on her face, trunk, and lower extremities.

Cases of patients with a “magenta colon” from radiologic barium have been reported every 3 to 4 years.²³ This condition can occur with either upper or lower barium contrast studies, resolves over 4 to 7 days, and can be present without visible residual barium. It was noted that most reactions are mild, but, occasionally, are severe enough to hamper determining whether there is true colitis. The severe reaction is characterized by edema, loss of all vascular markings, and redness to almost a magenta color, but without ulcerations, friability, necrosis, or exudate. Biopsies show inflammatory changes.

GENOTOXICITY

Barium Sulfate

The genotoxicity of barium sulfate was evaluated using murine fibroblasts in the *in vitro* single-cell gel (comet) assay. The fibroblasts were exposed for 5 h (at 37°C) to barium sulfate at final concentrations ranging from 10 to 1,000 µg/ml. Vehicle control cultures were exposed to phosphate-buffered solution, and positive control cultures were exposed to 10 µM hydrogen peroxide. A total of 50 randomly captured comets per treatment (25 cells from each slide) were examined using a fluorescence microscope. Barium sulfate did not increase cell mortality and was not genotoxic, i.e., did not induce DNA breakage. The positive control caused a significant increase ($P = 0.02$) in tail moment, when compared to the negative control.²⁴ Barium sulfate (1 to 1,000 µg/mL) also was not genotoxic in human peripheral blood lymphocytes in the *in vitro* single-cell (comet) assay.²⁵

SUMMARY

Barium Sulfate functions as an opacifying agent in cosmetic products. According to information supplied to the FDA by industry as part of the VCRP in 2014, barium sulfate is being used mostly in leave-on products. Results from a survey of ingredient use concentrations provided by the Council in 2013 indicate that this ingredient is being used at concentrations up to 0.99% and 37% in rinse-off and leave-on products, respectively.

The inorganic sulfates are typically manufactured by mining of natural minerals (as many inorganic sulfates occur naturally in hydrated form) or by reaction of available ore or inorganic oxides, hydroxides, or carbonates, with sulfuric acid. These methods produce hydrated inorganic sulfates. To produce the anhydrous salts, an additional step of dehydration (e.g., by heating and reduced pressure) must occur.

Chromium has been detected in commercial samples of pharmaceutical grade barium sulfate at concentrations ranging from 0.45 to 1.06 µg/g.

Rats inhaled barium sulfate (40 mg/m³) for up to 2 months, and barium was detected in the bone and lungs. Lymph transport was said to have been negligible. Following endotracheal administration of barium sulfate (BarosperseTM) into rats (0.25 ml volume) and dogs (1.75 ml/kg dose), the test material was virtually cleared from the trachea and stem bronchi in 30 minutes. Barium was detected in the alveoli at microscopic examination.

Groups of young male CBL-Wistar albino rats received a 150% w/v barium sulfate suspension orally at doses up to 375 g/kg. A mean LD₅₀ of 307 ± 29 g/kg was reported.

Following endotracheal administration of barium sulfate (BarosperseTM, 1.75 ml/kg dose), 2 of 3 dogs died during the first 24 h. Diffuse alveolarization of the test material was observed in both dogs.

Subpannicular injection of sterile barium sulfate in newborn rabbits produced an acute inflammatory response that was observed clearly at 24 h and well-established by 48 h. Following endotracheal administration of barium sulfate (BarosperseTM) into rats (0.25 ml volume) and dogs (1.75 ml/kg dose), a mild acute response and a chronic inflammatory response were observed. In another study, groups of Wistar rats were exposed repeatedly to barium sulfate dust for up to 90 days. Barium sulfate dust produced a low degree of inflammation at an exposure concentration of 75 mg/m³.

Barium sulfate dust, when inhaled, leads to a benign form of pneumoconiosis (baritosis), which occurs primarily in miners and workers in the lithopone industry.

Barium sulfate-granulomatosis of the lung, appendix, and other tissues has occurred during radiographic procedures. Additionally, though rare, hypersensitivity reactions have been reported after radiographic procedures that involve dosing with barium sulfate.

Barium sulfate was not genotoxic to murine fibroblasts or human peripheral blood lymphocytes in the *in vitro* comet assay.

DISCUSSION

The Panel noted that the history of safe medical use of barium sulfate indicates no significant toxicity concerns relating to systemic exposures to these ingredients. Furthermore, the extensive clinical experience of the Panel, including the results of numerous patch tests, indicates that barium salts do not have the potential to induce sensitization. The Panel noted that salts of sulfuric acid, such as sodium sulfate, can be irritating to the skin, so cosmetic products containing barium sulfate should be formulated to be non-irritating. Barium sulfate is being used in leave-on products at concentrations up to 37%.

Barium sulfate is being used at concentrations up to 15% in cosmetic products that are sprayed (perfumes) and at concentrations up to 15.8 % in powders (face and dusting powders). The Panel discussed the issue of incidental inhalation exposure from propellant and pump sprays and powders, and considered pertinent data indicating that incidental inhalation exposures to this ingredient in such cosmetic products would not cause adverse health effects. The data considered include data characterizing the potential for this ingredient to cause acute toxicity, inflammation, and genotoxicity. The Panel noted that 95% – 99% of droplets/particles produced in cosmetic aerosols would not be respirable to any appreciable amount. Coupled with the small actual exposure in the breathing zone and the concentrations at which the ingredients are used, the available information indicates that incidental inhalation would not be a significant route of exposure that might lead to local respiratory or systemic effects. A detailed discussion and summary of the Panel's approach to evaluating incidental inhalation exposures to ingredients in cosmetic products is available at <http://www.cir-safety.org/cir-findings>.

CONCLUSION

The CIR Expert Panel concluded that barium sulfate is safe in the present practices of use and concentration in cosmetics, when formulated to be non-irritating.

Table 1. Frequency and Concentration of Use According to Duration and Type of Exposure for Barium Sulfate.^{6,7}

	# of Uses	Conc. (%)
Totals/Conc. Range	398	0.001-37
Duration of Use		
<i>Leave-On</i>	385	0.001-37
<i>Rinse off</i>	4	0.0035-0.99
<i>Diluted for (bath) Use</i>	NR	0.94
Exposure Type		
<i>Eye Area</i>	98	0.01-18.6
<i>Incidental Ingestion</i>	87	0.04-37
<i>Incidental Inhalation-Sprays</i>	22	15
<i>Incidental Inhalation -Powders</i>	43	0.034-15.8
<i>Dermal Contact</i>	235	0.0035-20
<i>Deodorant (underarm)</i>	NR	NR
<i>Hair - Non-Coloring</i>	NR	0.55
<i>Hair-Coloring</i>	NR	0.62
<i>Nail</i>	69	0.001-3
<i>Mucous Membrane</i>	88	0.04-37
<i>Baby Products</i>	NR	NR

NR = Not Reported; NS = Not Surveyed; Totals = Rinse-off + Leave-on Product Uses.

Note: Because each ingredient may be used in cosmetics with multiple exposure types, the sum of all exposure type uses may not equal the sum total uses.

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